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A physical law describing the relationship of the measurable properties of an ideal gas, where P (pressure) \times V (volume) = n (number of moles) \times R (the gas constant) \times T (temperature in Kelvin). It is derived from a combination of the gas laws of Boyle, Charles, and Avogadro.

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Express your answer to three significant figures and include the appropriate units. Part A: Substance A: higher boiling point and has a higher heat of vaporization Substance B: Has weaker intermolecular forces and is a gas at 300 mmHg

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Answer Key 1303 Answer Key Chapter 1 1. Place a glass of water outside. It will freeze if the temperature is below 0 °C. 3. (a) law (states a consistently observed phenomenon, can be used for prediction); (b) theory (a widely accepted explanation of the behavior of matter); (c) hypothesis (a tentative explanation, can be investigated by experimentation) 5.

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